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*Chapter 12: Basic Review Worksheet 1. A chemical bond is a force that holds two or more atoms together and makes them function as a unit. The principal types of chemical bonding are ionic bonding, pure covalent bonding, and polar covalent bonding. 2. Ionic bonding results when elements of very different electronegativities react with each other.*

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Chapter 12: Standard Review Worksheet 1. The strength of a chemical bond is commonly described in terms of the bond energy, which is the quantity of energy required to break the bond. 2. We know that ionically bonded solids do not conduct electricity in the solid state (because

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Chapter 11: Basic Review Worksheet 1. What is electromagnetic radiation? 2. Sketch a representation of a wave and indicate on your drawing one wavelength of the wave. 3. Explain what it means for an atom to be in an excited state and what it means for an atom to be in its ground state. 4. What is a photon? 5. Describe Bohr's model of the ...

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C30 Ch12 Rev WS Chemistry 30 - Chapter 12 Review Worksheet Name: \_\_\_\_\_

1. Historically, the production of a small volume of pure metal from a larger volume of metal ore was called \_\_\_\_\_. The substance in the

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*reaction which*

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*is used to measure the pressure of samples of gas in the laboratory. A manometer consists basically of a U-shaped tube filled with mercury, with one arm of the*

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*reactant compounds for the ions that. combine to form the precipitate.*

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*being dispersed among water molecules. (See the discussion of question 2, Chapter 15: Basic Review Worksheet.) 3. To say that a solution is saturated does not necessarily mean that the solute is present at a high concentration. For example, magnesium hydroxide only dissolves to a very small*

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*Chapter 19: Challenge Review Worksheet 1. There are four stable isotopes of iron with mass numbers 54, 56, 57, and 58. There are two radioactive isotopes: iron-53 and iron-59. Predict modes of decay for each of these isotopes and explain. 2. Neon consists primarily of two isotopes with mass numbers 20 and 22, with a small amount of a*

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