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guide True/False  
Indicate whether the  
statement is true or*

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**false. \_\_\_\_\_ 1. The actual yield is always lower than the theoretical yield. Matching Match each item with the correct statement below. ? a. stoichiometry ? ...**

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Chapter 11  
Stoichiometry | pdf  
Book Manual ...**

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**Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and**



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***the amounts of products formed by a chemical reaction is called stoichiometry.***

***Stoichiometry is based on the law of conservation of mass,***

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***Chapter 11:  
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Stoichiometry  
Section 11.1 What is  
stoichiometry? In***

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***your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false.***

***\_\_\_\_\_ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by***

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**a chemical reaction  
is called  
stoichiometry.**

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stoichiometry  
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In Section 11.3 , for  
example, you  
learned how to**

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***express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium***

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*dichromate  
produces 4 mol of  
water). This section*

...

## **VIBRATIONS AND WAVES**

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Chapter 11 –  
Stoichiometry  
(Rough outline of  
the chapter, please  
use the book, notes  
& homework to***

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*study.*) **11.1 Defining  
Stoichiometry  
Vocab • Key**

***stoichiometry •  
mole ratio Concepts  
Using Balanced  
Equations • Number  
of Atoms • Number  
of Molecules •  
Number of Moles •  
Mass o Law of  
Conservation of  
Mass • Volume.***



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Stoichiometry  
Vocab •**

**stoichiometry •**

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***mole ratio Concepts***

***Using Balanced***

***Equations • Number***

***of Atoms • Number***

***of Molecules •***

***Number of Moles •***

***Mass o Law of***

***Conservation of***

***Mass • Volume 11.2***

***...***

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***Guide Stoichiometry***

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**11.1 Defining**

**Stoichiometry MAIN**

**Idea** *The amount of each reactant*

*present at the start of a chemical*

*reaction determines how much product*

*can form. 11.2*

**Stoichiometric**

**Calculations MAIN**

**Idea** *The solution to every stoichiometric*

*problem requires a*

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**balanced chemical  
equation. 11.3**

**Limiting Reactants  
MAIN Idea A  
chemical reaction**

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companion course  
helps students learn  
the essential  
chemistry lessons  
of products and**

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*reactants.*  
Stoichiometry

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**11 YCHAPTER**

**Section 11.2**

**Stoichiometric**

**Calculations In your  
textbook, read about  
mole-to-mole**

**conversion. Read  
the following**

**passage and then**



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***solve the problems.***

***In the equation that***

***follows...Key***

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and Change Chapter  
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***CHAPTER Section***

***11.1 continued In***

Read Online  
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*your textbook, read  
about mole ratios.*

*Answer the  
questions about the  
following chemical  
reaction. sodium +  
iron(III) oxide  $\rightarrow$   
sodium oxide + iron  
 $6\text{Na}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow 3\text{Na}_2\text{O}(s) + 2\text{Fe}(s)$*

*15. What is a mole  
ratio? ... Study  
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**11.3 Limiting**

**Reactants Date**

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***In your textbook,  
read about why  
reactions stop and  
how to determine  
the limiting reactant.  
Study the diagram  
showing a chemical***

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***reaction and the  
chemical equation  
that represents the  
reaction. Then  
complete the table.  
Show your  
calculations for  
questions 25—27***

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12: Stoichiometry .  
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**Name CHAPTER**

**Section 11.2**

**continued Date**

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**In your textbook,**

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***read about mole-to-mass and mass-to-mass conversions. Solving a mass-to-mass problem requires the four steps listed below.***

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