

## Chapter 11 Review Gases Section 3 Short Answer

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explains some of the properties of ideal gases. In this chapter, you will study the predictions of kinetic-molecular theory for gases in more detail. This includes the relationship among the temperature, pressure, volume, and amount of gas in a sample. SECTION 11.1 VOCABULARY pressure newton barometer millimeters of mercury atmosphere of ...

Chapter 11 Review Gases Section

CHAPTER 11 REVIEW Gases SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. b Pressure surf f a o c r e ce area. For a constant force, when the surface area is tripled the pressure is (a) doubled. (b) a third as much. (c) tripled. (d) unchanged.

Dodd, Mead and Company

CHAPTER 11 REVIEW Molecular Composition of Gases MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. The average speed of a gas molecule is most directly related to the . (a) polarity of the molecule (b) pressure of the gas (c) temperature of the gas (d) number of moles in the sample 2.

CHAPTER 11 REVIEW Gases - Manasquan Public Schools

Section Goals and Introductions Section 11.1 Gases and Their Properties Goals To describe the particle nature of both real and ideal gases. To describe the properties of gases that can be used to explain their characteristics: volume, number of particles, temperature, and pressure.

11 Molecular Composition of Gases - Madison Public Schools

Chapter 11 Section 1 Gases and Pressure •Torricelli reasoned that if the maximum height of a water column depended on its weight, then mercury, which is about 14 times as dense as water, could be

Gases - Los Angeles County High School for the Arts

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CHAPTER 11 Gases - St. Charles Parish

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Chapter 11 Gases - An Introduction to Chemistry

Modern Chemistry 99 Gases CHAPTER 11 REVIEW Gases SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_ List the following gases in order of increasing rate of effusion. (Assume all gases are at the same temperature and pressure.) (a) He (b) Xe (c) HCl (d) Cl<sub>2</sub>

The Stationery Office

CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L.

Chapter 11 Section 1 Gases and Pressure Objectives

This video explains the concepts from your packet on Chapter 11 (Liquids and Intermolecular Forces), which can be found here: <https://goo.gl/UhCv2b> Section 11.1: A Molecular Comparison of Gases ...

Chapter 11 Liquids and Intermolecular Forces

View Homework Help - chapter 11 section 1 from CHEM intro at Wenatchee High School. W Gases AK /} SHORT ANSWER Answer the following questions in the space provided. I 1- & Pressure = wL. For a

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11.1 Gases and Their Properties 463 For an ideal gas (in which the particles occupy no volume and experience no attractions or repulsions), gas pressure and volume are inversely proportional.

Section 11.3 Review Flashcards | Quizlet

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Ex C pg 370 A sample of oxygen gas has a volume of 150.0 mL when its pressure is 0.947 atm. What will the volume of the gas be at a pressure of 0.987 atm if the

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Chemistry Chapter 11 Gases. STUDY. PLAY. Pressure. The amount of force exerted per unit area of a surface. Newton. The SI unit for force; the force that will increase the speed of a 1 kg mass by 1 m/s each second that the force is applied. Barometer. An instrument that measures atmospheric pressure.

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