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CHAPTER . 11 . REVIEW . Gases . SHORT ANSWER . Answer the following questions in the space provided. 1. Consider the following data table: Approximate pressure (kPa) Altitude above sea level (km) 100 . o (sea level) 50 5.5 (peak of Mt. Kilimanjaro) 25 11 Get cruising altitude} < 0.1 22 (ozone layer) ~.

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Unformatted text preview: Name w Date W Class CHAPTER 11 REVIEW ' Gases SHORT ANSWER Answer the following questions in the space provided. force WW.._For a constant force, when the surface area is tripled the surface area 'I . b Pressure 1 pressure is (a) doubled. (b) a third as much.

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Chapter 11 - Gases Chapter 11 focuses on gas behavior and the gas

laws. In Chapter 10, students were given an overview of the kinetic-molecular theory of matter and discussed how this theory explains the chemistry of particles in the solid, liquid, and gas phases.

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Temperature of gas depends on average kinetic energy of particles Ch 10 pg 329-330 Parameters for an ideal gas Kinetic molecular theory only applies to "ideal" gases. Most gases are only nearly ideal if pressure is not too high and temperature is not too low

Chapter 11 - Gases

Review Previous Concepts Gases CHAPTER 11 1. What is a mole? What information does the molar mass of a substance relate about the substance? 2. What is the kinetic-molecular theory of matter? Key Concepts LOUISIANA STANDARDS LA.PS.29 Predict the properties of a gas based on gas laws. (e.g., temperature, pressure, volume). (PS-H-C7) 340 CHAPTER 11

Chapter 11 Review Gases Mixed Answer Key

CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances:

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Gases - Los Angeles County High School for the Arts

CHAPTER 11 REVIEW Gases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. c The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided by the mass of 1 mol. (d) divided by 22.4 L. 2. c For the expression $V = nRT/P$, which of the following will cause the volume to increase?

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Modern Chemistry 95 Gases CHAPTER 11 REVIEW Gases SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. State whether the pressure of a fixed mass of gas will increase, decrease, or stay the same in the following circumstances: _____ a. temperature increases, volume stays the same

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