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Reactions ...

11.2 Types of

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Reactions > 13 A

decomposition

reaction is a

chemical change

in which a

single compound

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*breaks down into
two or more*

*simpler
products. •*

*Decomposition
reactions*

*involve only one
reactant and two
or more*

*products. • The
products can be
any combination
of elements and
compounds. •*

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*Most
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reactions
require ...*

*Chapter 11:
Chemical
Reactions*

*A chemical
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reaction are the
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Reactions
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14.6 Energy
Balances for
Reactions.....

14-10 14.6.1

Example. Heat of
reaction for CO
 $+ \text{H}_2\text{O} = \text{CO}_2 +$
 H_2 14-10

14.6.2 Example.

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Reactions. First
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use coefficients
to balanced the
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conservation of
mass.

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11.1 -

Describing

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products on the
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that separates
them is called
yield. Reactants
Products.*

*Symbols in
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*chemistry study
chapter 11*

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Flashcards ...

Combination

Reactions The

first type of

reaction is the

combination, or

synthesis,

reaction. A

combination

reaction is a

chemical change

in which two or

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more substances
react to form a
single new
substance. As
shown in *Figure*
11.5, magnesium
metal and oxygen
gas combine to
form the
compound
magnesium oxide.

Chapter 11.2

Types of

Page 18/42

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*Flashcards /
Quizlet*

*A representation
of a chemical
reaction...*

*Reactants on the
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Reactions

Workbook Ans

substance that speeds up the reaction but is not used up in...

Whole numbers that are placed in front of the formulas in an e... Chemical

Equation A representation of a chemical reaction...

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CHAPTER 11

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old paradigms.
NOW is the time
to make today
the first day of
the rest of your
life.*

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Page 24/42*

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11.2 Types of

Chemical

Reactions > 29

Whether one

metal will

displace another

metal from a

compound depends

upon the

relative

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11.1 Describing

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Reactions 11

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Chemistry

Chapter 11. A

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11.2 Types of
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and combustion.

a chemical

change in which

a single

compound breaks

down into two or

more simpler

products; only

one reactant and

two or more

products.

Chapter 11.2

Types of

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with valuable
instructions,
information and
warnings.*

SECTION 11.1

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DESCRIBING
CHEMICAL
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321-329)

CHAPTER 11

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and the single

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Chapter 7 57 11.

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*Figure 11.2
Three common
chemical
reactions are
shown. When
methane gas
burns, it
combines with
oxygen to form*

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*carbon dioxide
and water. Iron
turns to red-
brown rust
(iron(III)
oxide) in the
presence of
oxygen in the
air. Hydrogen
peroxide
decomposes to
water and oxygen
when used as an
antiseptic.*

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CHAPTER 11,

Chemical Reactions (continued) 9.

Use the symbols

in Table 11.1 on

Page 38/42

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page 323 to
write a skeleton
equation for the
following
chemical
reaction.

Hydrochloric
acid reacts with
zinc to produce
aqueous zinc(II)
chloride and
hydrogen gas.

Balancing
Chemical

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Equations (pages
324-328) 10.
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and Study
Workbook ...
Every minute of
the day,
chemical.
reactions take
place—both
inside you and.*

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around you.

•After a meal, a series of chemical reactions. take place as your body digests food. •Plants use sunlight to drive the. photosynthetic processes needed to produce. plant growth.

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