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Balancing Equations: Practice Problems

A balanced chemical equation gives the number and type of atoms participating in a reaction,

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the reactants, products, and direction of the reaction. Balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow.

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Balancing Chemical Equations

Balancing Equations Race. 1)

_ C₃H₈ + O₂ ^ CO₂ + H₂O. 2) Al +

FeS₂ ^ Al₂O₃ + Fe

3) Na + Cl₂ ^

NaCl 4) H₂O₂ ^ H₂O + O₂. 5)

C₆H₁₂O₆ + O₂ ^ H₂O + CO₂ 6)

H₂O + CO₂ ^ C₇H₈ + O₂. 7)

NaClO₂ ^ NaCl + O₂. 8)

(NH₄)₃PO₄ + Pb(NO₃)₂ ^ Pb₃(PO₄)₂

+ NH₄NO₃. 9) BF₃ + Li₂SO₃ ^

6LiF + B₂(SO₃)₃ + Li₂O 10) C₇H₁₆ + O₂ ^

CO₂ + H₂O. 11) CaCO₃ + H₃PO₄

-> Ca₃(PO₄)₂ + H₂CO₃.

*Student Exploration- Balancing
Chemical Equations (ANSWER
KEY)*

Balancing Chemical Equations

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equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$ 6) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$

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Balancing from balancing chemical equations practice worksheet answer key ,
source:acthum.net At times, the total amount of energy involved with a reaction could be so small that the change in energy isn't readily noticeable.*

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1 (1) - Balancing Chemical ...

To set up an equation in the Chemical Equations Gizmo, type the chemical formulas into the text boxes of the Gizmo. First, type in "H₂+O₂" in the Reactants box and "H₂O" in the Products box.

49 Balancing Chemical Equations Worksheets [with Answers]

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Answer Key Vocabulary:
coefficient, combination,
compound, decomposition,
double replacement, element,
molecule, product, reactant,
single replacement, subscript
Prior Knowledge Questions (Do
these BEFORE using the Gizmo.)
[Note: The purpose of these
questions is to activate prior
knowledge and get students
thinking.*

*Balancing Equations Practice
Problems Worksheet Answers Key
Balancing Chemical Equations
Worksheet 1 Answer Key -*

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Because they are advised to accomplish this some students don't bother and just balance it, but a few of them try being logical and wish to understand the real reason. As they are eager to learn what the scene will be my pupils are engaged in the activity and the way in which the math problems will be nestled within.

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*Balancing Equations: Answers to
Practice Problems 1. Balanced
equations. (Coefficients equal to
one (1) do not need to be shown
in your answers).*

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How to Balance Equations - Printable Worksheets

The first step to balance the equation is to write down the chemical formula of reactants that are listed on the left side of the chemical equation. After this, you can list down the products on the right hand side of the chemical equation. There is an arrow between the sides, signaling the direction the reaction is happening in.

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Balancing Chemical Equations -
Answer Key Balance the

equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightleftharpoons 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightleftharpoons 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightleftharpoons 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightleftharpoons 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightleftharpoons 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$ 6) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightleftharpoons 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$ 7) $1 \text{ CH}_4 + 2 \text{ O}_2 \rightleftharpoons 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$ 8) $1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightleftharpoons 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$ 9) $2 \text{ C}_8\text{H}_{18} + 25 \text{ O}_2 \rightleftharpoons 16 \text{ CO}_2 + 18 \text{ H}_2\text{O}$ 10) $1 \text{ FeCl}_3 + 3 \text{ NaOH} \rightleftharpoons 1 \text{ Fe(OH)}_3 + 3 \text{ NaCl}$ 11) $4 \text{ P} + 5 \text{ O}_2 \rightleftharpoons 2 \text{ P}_2\text{O}_5$

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