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Hybridization of ClF₃: Hybridization of Cl in Chlorine ...

Arrhenius acid in the aqueous solution increases the concentration of protons or H⁺ ions. For example, hydrochloric acid in the water. HCl undergoes dissociation reaction to produce H⁺ ion and Cl⁻ ion, as explained below. The concentration of the H⁺ ions is increased by forming hydronium ion. $\text{HCl (aq)} \rightarrow \text{H}^+ \text{ (aq)} + \text{Cl}^- \text{ (aq)}$

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Note:-All materials will be revised shortly for the session 2018-19. SECONDARY CLASSES 2017-18 Maths Class X Question Bank for 2017-18 Physics Reference Study Material Class X for 2017-18 Chemistry Reference Study Material Class X for 2017-18 Biology Reference Study Material Class X for 2017-18 Maths Class IX Question Bank for 2017-18 Physics Reference Study Material...

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The central atom Cl needs three unpaired electrons to bond with three F-atoms. ClF₃ should consist of 3 bond-pairs and 2 lone-pairs. One 3s, three 3p and one of the 3d orbitals of Cl participate in the hybridization and five sp³d hybrid orbitals are formed.

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Solution: Question 2. In a reaction, $2A \rightarrow \text{products}$, the concentration of A decreases from 0.5 mol L⁻¹ to 0.4 mol L⁻¹ in 10 minutes. Calculate the rate during this interval. Solution: Question 3. For a reaction, $A + B \rightarrow \text{product}$; the rate law is given by, $r = k[A]^{1/2} [B]^2$. What is the order of the reaction? Solution: Question 4.

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Solution: First of all, the compound has complex positive part $[\text{Co}(\text{H}_2\text{O})_4(\text{NH}_3)_2]^{3+}$ therefore, according to IUPAC conventions, positive part will be named first. Secondly, in writing name of ...

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Arrhenius theory - Arrhenius acid and base, examples

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Sulphur . Selenium . Tellurium. Sulphurous acid H₂SO₃. Sulphuric acid H₂SO₄. Peroxomonosulphuric acid H₂SO₅ (Caro's acid). Peroxodisulphuric acid. H₂S₂O₈ (Marshall's acid). Thio sulphuric acid H₂S₂O₃. Dithionic acid H₂S₂O₆. Pyrosulphuric acid H₂S₂O₇. Selenious acid H₂SeO₃. Selenenic acid H₂SeO₄. Tellurous acid H₂TeO₃. Telluric acid H₂TeO₄.

NCERT Solutions for Class 10 Science Chapter 5 Periodic ...

Example, NH_4OH and NH_4Cl . $\text{pH} = \text{pK}_b + \log \frac{[\text{Conjugated acid}]}{[\text{Base}]}$ Solubility and Solubility Product : 1. Amount (moles) of the salt that has made the solution saturated per liter of solution is called the solubility of the salt. 2. For salt AB. $\text{AB} \leftrightarrow \text{A}^+ + \text{B}^-$ $K_{sp} = [\text{A}^+][\text{B}^-]$ 3.

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