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Answered: You want to prepare 450.0 mL of a 145... | bartleby
Non-metals can react with salt solution; the more reactive element will displace the less reactive non-metal. $2 \text{NaBr (aq)} + \text{Cl}_2(\text{aq}) \rightarrow 2\text{NaCl} \dots$ At anode $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$ Refining of Metals. ... Refer to our NCERT Solutions and prepare answers for every question in the chapter. Use flashcards, mind maps, practice worksheets to memorise the ...

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You want to prepare 450.0 mL of a 145 ppm K^+ ion solution. The density of the solution is 1.01 g/mL. You have at your disposal a solution that is 0.728 M in K_2SO_4 . How many mL of the K_2SO_4

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solution do you need to use? 1.16
mL 3.75×10^{-6} mL 2.56 mL 0.860
mL 116 mL

Revision Notes for Science
Chapter 3 - Metals and Non ... -
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question on Electrostatics for
JEE, NEET & CBSE

A: Given the molarity of lipid
solution, $C = 0.11 \text{ mM}$ Volume of
lipid solution, $V = 40 \mu\text{L}$

Required... question_answer Q:

Name the following ionic

compounds: 1) $\text{V}(\text{SO}_4)_3$ 2)

Zn_3P_2 3) $\text{Sr}(\text{C}_2\text{H}_3\text{O}_2)_2$ 4) Cu_2O

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