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6 POGIL □ Activities for High School Chemistry 18. Based on the data in Model 3, does the equilibrium ratio of product to reactant depend on the percent of the molecules that reacted in the forward and reverse reactions? If yes, describe the relationship. 19. Predict the final product to reactant ratio for the reaction in Model 1 if the initial concentrations of A and B are 15.0 moles and 5 ...

Equilibrium.pdf - Equilibrium At what point is a ...

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Answers to Student Questions (from the articles) 6. ChemMatters Puzzle: Fluster for Chemists 11 ... ($3 \times 1.3806 \times 10^{-23} \text{ J/K} \times 1 \text{ kg-m/s}^2 \times 273 \text{ K} \times 1 \text{ mole} \times 6.022 \times 10^{23} \text{ molecule} \times 1000 \text{ g} \dots$ A student could also design a demonstration showing that the surface-volume ratio of an object affects the rate at which heat is lost by any of the above ...

ChemMatters Teacher's Guide - American Chemical Society

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classifying chemical reactions worksheet answers chapter 21, ordinary chemical reactions, positive ions are produced by removing electrons from the atoms. Li^+ $3+$ In the pictures below draw in the number of electrons needed to make the ion named in the box. O^{2-} $8+$ F^{-} $9+$ **NEGATIVE IONS** Negative ions have more electrons than protons. Since the number of protons is unchanged from the neutral

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