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**Answers**

In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by



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comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation, just as we did in Section 11.4 .

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\_\_\_\_\_21. Metal X replaces the ions of metal Y from solution, but it cannot replace the ions of metal Z from solution. The order these metals should have in the activity series (from top to bottom) is a. x, y, z b. z, x, y c. y, x, z d. z, y, x \_\_\_\_\_22.



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quantities in chemical reactions is a  
subject of chemistry. Mole ratio.

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Name \_\_\_\_\_ Date \_\_\_\_\_ Class **STOICHIOMETRY**  
Chapter Test A A. Matching Match each description in Column B with the correct term in Column A. Write the letter of the correct description on the line. Column A  
1. actual yield 2. limiting reagent 3.  
theoretical yield 4. stoichiometry 5.

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percent yield 6.

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Chapter 12 Multiple Choice Identify the letter of the choice that best completes the statement or answers the question. \_\_\_\_\_ 1.  
The first step in most stoichiometry

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problems is to \_\_\_\_\_. a. add the coefficients of the reagents c. convert given quantities to volumes b. convert given quantities to moles d. convert given quantities to masses \_\_\_\_ 2.

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You can skip questions if you would like  
and come back to them ...

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d. 12% e. 2% 11. How many grams of  $\text{H}_2\text{O}$  will be formed when 32.0 g  $\text{H}_2$  is allowed to react with 16.0 g  $\text{O}_2$  according to  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$  a. 9.00 g b. 16.0 g c. 18.0 g d. 32.0 g e. 36.0 g 12. When 2.00 g of  $\text{H}_2$  reacts with 32.0 g of  $\text{O}_2$  in an explosion, the final gas mixture will contain: a.  $\text{H}_2$ ,  $\text{H}_2\text{O}$ , and  $\text{O}_2$  b.  $\text{H}_2$  and

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H<sub>2</sub>O only c. O<sub>2</sub> ...

**Chapter 12.2: Stoichiometry of  
Reactions in Solution ...**

Stoichiometry Chapter 3! Stoichiometry:  
Calculations with Chemical Formulas and  
Equations. Stoichiometry Anatomy of a  
Chemical Equation  $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g})$

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$\text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{g})$  Stoichiometry ... •  
1/12 mass of the  $^{12}\text{C}$  isotope.

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